Common Polyatomic Ions					
Charge	**Name**	**Formula**	**Charge**	**Name**	**Formula**
1+	ammonium	NH_4^+	1−	perchlorate	ClO_4^-
1−	acetate	$CH_3CO_2^-$	1−	periodate	IO_4^-
1−	bromate	BrO_3^-	1−	permanganate	MnO_4^-
1−	chlorate	ClO_3^-	2−	carbonate	CO_3^{-2}
1−	chlorite	ClO_2^-	2−	chromate	CrO_4^{-2}
1−	cyanide	CN^-	2−	dichromate	$Cr_2O_7^{-2}$
1−	bicarbonate (hydrogen carbonate)	HCO_3^-	2−	oxalate	$C_2O_4^{-2}$
1−	hydrogen sulfate	HSO_4^-	2−	sulfate	SO_4^{-2}
1−	hydroxide	OH^-	2−	sulfite	SO_3^{-2}
1−	hypochlorite	ClO^-	2−	thiosulfate	$S_2O_3^{-2}$
1−	iodate	IO_3^-	3−	arsenate	AsO_4^{-3}
1−	nitrate	NO_3^-	3−	borate	BO_3^{-3}
1−	nitrite	NO_2^-	3−	phosphate	PO_4^{-3}

Nomenclature Rules Summary:

1. Acids
   a. Binary acid: hydro_____ic acid
   b. Oxy acid:
      i. hypo___ite — hypo___ous acid
      ii. -ite — ___ous acid
      iii. -ate: — _____ic acid
      iv. per___ate — per___ic acid

2. Covalent — (num. prefix)_____ (num. prefix)_____ide
   (if first number prefix is "mono", drop it)

3. Ionic
   a. Binary compound — _____ _____ide
   b. Polyatomic ion — _____ name from chart
   c. Metal binary — _____(NUM) _____ide
      Metal & polyatomic — _____(NUM) name from chart
      (NUM) = valence charge of metal